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Molecular mass of $\text{KClO}_3 = 39 + 35.5 + 3 \times 16 = 122.5$
 $39 + 35.5 + 3 \times 16 = 122.5$. No. of mole of $\text{KClO}_3 = \frac{245 \text{ g}}{122.5 \text{ g}} = 2$ mole. 2 mole of KClO_3 contains. 2 mole of K = $2 \times 6.022 \times 10^{23} = 12.044 \times 10^{23}$
 $2 \times 6.022 \times 10^{23} = 12.044 \times 10^{23}$ K atoms. 2 mole of Cl = $2 \times 6.022 \times 10^{23} = 12.044 \times 10^{23}$
 $2 \times 6.022 \times 10^{23} = 12.044 \times 10^{23}$ Cl atoms.

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SOME BASIC CONCEPTS OF CHEMISTRY

Chemistry is the science of molecules and their transformations. It is the science not so much of the one hundred elements but of the infinite variety of molecules that may be built from them. Roald Hoffmann Science can be viewed as a continuing human effort to systematise knowledge for describing and understanding nature.

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$= 1 \text{ kg m}^{-2} \text{ kg m}^{-2} \text{ s}^{-2} \text{ s}^{-2} \text{ Pa} = 1 \text{ kg m}^{-1} \text{ kg m}^{-1} \text{ s}^{-2} \text{ s}^{-2}$
 $\therefore \therefore \therefore \text{Pressure (P)} = 1.01332 \times 10^5 \text{ Pa}$ Q14.

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Number of moles of solute (KCl) = given mass/ molecular mass. Number of moles of solute (KCl) = 7.45 g/ 74.5 g mol⁻¹ = 0.1 mol. Molality = Number of moles of solute/Mass of solvent in kg. Molality = 0.1 mol /0.1 kg = 1 mol kg⁻¹. Ans: The molality of solution is 1 mol kg⁻¹ or 1 m.

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Molality, Molarity, Mole fraction: Numerical problems

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Chapter 1 - Some Basic Concepts of Chemistry - BYJUS

Concrete numbers and base units. Many parameters and measurements in the physical sciences and engineering are expressed as a concrete number—a numerical quantity and a corresponding dimensional unit. Often a quantity is expressed in terms of several other quantities; for example, speed is a combination of length and time, e.g. 60 kilometres per hour or 1.4 kilometres per second.

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